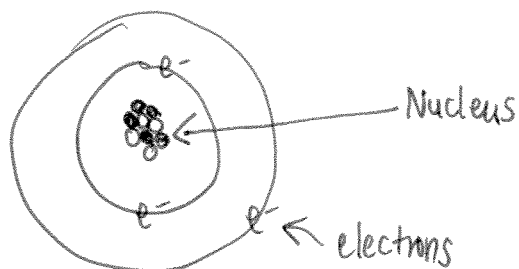


33. Why did the model of the atom change over time?

New experiments gave evidence that didn't fit current models so the models changed to reflect the new info.

34. Draw and describe the simple atomic model. What are the charges of the subatomic particles?

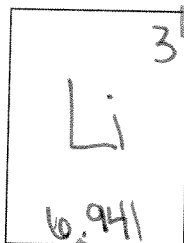


protons: +
neutrons: 0
electrons: -

LESSON 12: Atomic Number and Atomic Mass

35. Atomic number is the # of protons in an atom and determines an atom's identity.

36.



a. Fill in this block, showing the information from the periodic table for lithium.

b. How many protons are in an atom of lithium? 3

c. What is lithium's atomic number? 3

Round to be protons + neutrons

d. How many protons, neutrons, and electrons are in a neutral lithium atom?

1) Atomic # = 3 = 3 protons

2) Mass # = 7 so 7 - 3 = 4 neutrons

3) Neutral atom has 3 electrons to balance out the protons

37. Complete the following table using a periodic table.

Element	Symbol	Atomic Number	Number of Protons	Number of Electrons	Number of Neutrons	Average Atomic Mass
vanadium	V	23	23	23	(51-23=28) 28	50.942
chlorine ion	Cl ⁻	17	17	18	(35-17=18) 18	35.453
Phosphorus	P	15	15	15	(31-15=16) 16	30.974